



# Infrared Spectroscopy Study of Pure Oxides PbO, SnO<sub>2</sub> and their Comparison

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## Abstract

Nanocrystals of semiconducting metal oxides have attracted great interest due to their interesting properties. Therefore, our study came to some physical properties of pure tin oxide and pure lead oxide. Infrared spectral measurements of pure tin and lead oxides showed the presence of vibrational frequencies containing five frequencies for tin oxide and nine vibrational frequencies for lead oxide. The permeability spectra of the SnO<sub>2</sub> and PbO compounds were measured, and we calculated each of  $n$ ,  $A$ ,  $\alpha$ ,  $\sigma_{opt}$ , and it was found through the comparison of these amounts for the two compounds that the optical conductivity =  $1.214687 (\Omega\text{cm})^{-1}$ ,  $n=9.949$ ,  $A=2$ ,  $\alpha=46 \text{ cm}^{-1}$ , All previous values are in agreement with the same vibrational frequency of  $\nu = 3450 \text{ cm}^{-1}$  for the lead oxide compound, which is the highest possible.

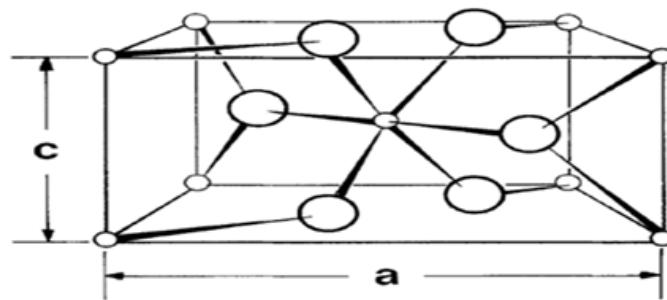
**Keywords:** Tin oxide; Lead oxide; Infrared spectrum; Absorbance; Absorption coefficient; Refractive index; Optical conductivity

## Introduction

Transparent conducting oxides are technologically important materials with a wide range of applications in electronics, solar cells, flat panel displays, and more. Where these materials combine a unique property of being optically transparent and electrically conductive, and among these materials [1]. Where tin oxide (SnO<sub>2</sub>) is one of the most important transparent conducting oxides (TCOs) due to its unique optical and electrical properties, in addition to being thermally and chemically stable, and is considered a material with high mechanical specifications as it has a low specific electrical resistance and optical transmittance. High in the visual field and stability when heat treatment.

It is among the n-type semiconductor materials with an ener

gy interval of 3.6 eV. The electrical conductivity of the oxide comes from the unevenness in the structure of these materials due to the oxygen voids. Conductivity can be improved by adding impurities. Electrical conductivity is also affected by heat [1]. The primary cell of tin oxide contains six atoms, two tin atoms and four oxygen atoms, as shown (Figure 1). Whereas, the tin atom Sn<sup>4+</sup> located in the center is surrounded by six oxygen atoms O<sup>2-</sup> forming a regular octahedron, while it surrounds all three atoms located on the vertices of an isosceles triangle, and it crystallizes according to the structure of the tetragonal, where  $a$ ,  $c$  are the network constants It is equal to ( $a = b = 4.737 \text{ \AA}$  and  $c = 3.185 \text{ \AA}$ ) and belongs to the  $mm/2P4$  space group (the quadrupole class rhomboid according to the preferred orientation (200) [2] (Figure 1).

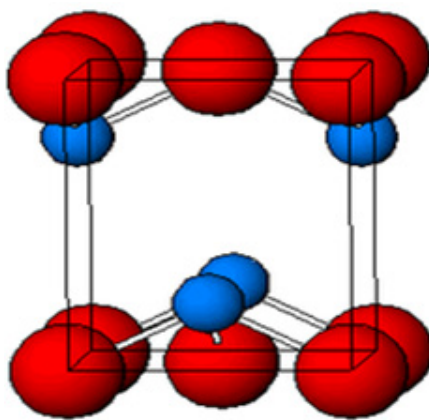


**Figure 1:** The crystal structure of tin oxide. The small circles indicate the locations of the tin atoms, the large circles indicate the locations of the oxygen atoms [2].

This oxide is characterized by high electrical conductivity in the visible field and reflective in the infrared field, where tin oxide has an electrical conductivity located in the range  $(10^3-10^8) (\Omega\text{cm})^{-1}$  at room temperature. This conductivity is affected by many factors, including heat, light and magnetic field in the presence of tiny amounts of impurity atoms, which makes it a very important material in electronic applications. Electrical conductivity is also considered one of the most important characteristics by which solids can be classified. Because of the wide forbidden field, an absorption edge appears in the ultraviolet range,  $\mu$  0.36, which makes this compound transparent to visible and near infrared rays [3]. One of these oxides is lead oxide, PbO, which is technologically important due to its attractive properties, as lead oxide has multiple levels of

reflectivity, which makes it suitable for optical storage devices [4].

Among the various properties that lead oxide possesses, PbO is an attractive material with low conductivity, interesting semiconductor and photoconducting properties which makes it suitable in laser technology and imaging device applications [5], as well as in applications such as gas sensors for CO<sub>2</sub>, also considered to be high-index materials, and PbO films were used as electrode materials for secondary lithium batteries [6]. The preparation of the PbO film is often complicated by its high volatility at low temperatures. At a low temperature, PbO is in the PbO- $\alpha$  tetragonal phase, and at a high temperature, it is in the axial phase. At 490 C°, PbO- $\alpha$  transition Phase to PbO- $\beta$  (Figure 2).



**Figure 2:** Crystal structure of lead oxide. The blue circles indicate the positions of the lead atoms, the red circles indicate the locations of the oxygen atoms [6].

## Infrared spectroscopy

Infrared spectroscopy is one of the branches of spectroscopy that studies the infrared region of the electromagnetic spectrum and includes a set of implementations, the most famous of which is absorption spectroscopy, which is used to determine the active groups in the compounds under study. Infrared spectroscopy is one of the basic methods in the study of materials, as it enables us to identify the structure of the material without affecting its properties. It relies on studying the spectra absorbed by the sample in the

energy range  $(14000-20) \text{ cm}^{-1}$ . The energy of red radiation is not sufficient to cause electronic excitations in most materials, but it is sufficient to cause stretching and bending vibrations in the bonds. And all types of bonds respond to this amount of energy to cause vibrations of this kind, so they absorb in the infrared region, provided that the absorption leads to a change in the dipole moment of the molecule, and these vibrations are quantized, and their occurrence means that the compound absorbs infrared energy in a specific part of the spectrum. Most of the spectral analyzes take place in the mid-infrared region  $(200-4000) \text{ cm}^{-1}$  where most of the molecular

vibrations occur and the researcher finds enough information to determine the molecular structure of the studied compounds .

### Principle of infrared spectroscopy

Natural particles vibrate according to all their vibration patterns, but with very weak amplitudes, just as the photon has a sinusoidal electric field if the frequency of the incoming photon corre-

sponds to the frequency of vibration of the regular patterns of the molecule, then the molecule will enter into resonance and vibrate with very large amplitudes, or in other words it absorbs the photon whose energy is equal to the energy necessary for the molecule to move from a low energy state to an excited state, and its energy is transformed into vibration energy as (Figure 3).

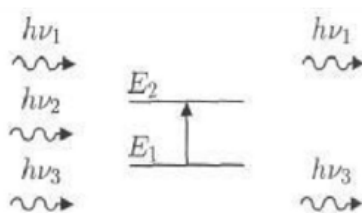


Figure 3: Infrared absorption.

Only a photon whose energy is  $2h\nu$  equal to the transition energy ( $E_2 - E_1$ ) is absorbed, thus creating an imbalance in the back-scattered radiation. The absorption of some incoming photons leads to the appearance of harmonic lines of photons that have not been bounced and do not appear in the infrared spectrum curve of the molecule, this absorption characterizes the bonds between the atoms, and since each vibration pattern corresponds to a single movement of the molecule, this means that there is a direct correspondence between the frequency of the absorbed radiation and the structure of the molecule. When the infrared rays are returned to a medium, they suffer reflection, absorption and permeation from the medium. The IR spectrum of the pure oxides ZnO and SnO<sub>2</sub> was measured using the aforementioned spectrometer. (Figure 3) shows the IR spectrum of those samples.

#### We Measured

**1. Transmittance T:** It is defined as the ratio between the intensity of the transmitted radiation to the intensity of the incoming radiation. It was taken from the device itself, and then using the appropriate mathematical relations, other optical coefficients were calculated.

**2. Absorbance A:** It is the ratio between the intensity of the absorbed radiation to the intensity of the incident radiation, and it is calculated from the equation: [7]

$$A = \log\left(\frac{100}{T\%}\right) \quad (1)$$

Where: T represents the Transmittance.

**3. Reflectivity R:** It is the ratio between the intensity of the reflected radiation to the intensity of the incident radiation. According to the law of conservation of energy, R can be calculated from the equation: [8]

$$A + T + R = 1 \quad (2)$$

**4. Absorption coefficient  $\alpha$ :** It is defined as the ratio between the decrease in the energy of the incoming radiation per unit distance in the direction of propagation of the incident light wave inside the medium, and it is calculated from the equation: [9]

$$\alpha = 2.303 \frac{A}{d} \quad (3)$$

Where:  $d=1$  mm is the material thickness.

**5. The damping factor K:** It is defined as the amount of energy absorbed by the atoms of the studied material from the energy of the incident radiation photons, and it is calculated from the equation [10]

$$K = \frac{\alpha}{4\pi\nu} \quad (4)$$

where  $\alpha$  is the absorption coefficient.

**6. Refractive index n:** It is the ratio between the speed of light in a vacuum to its speed in the medium, and it is calculated from the equation: [11]

$$n = \left[ \left( \frac{1+R}{1-R} \right)^2 - (K^2 + 1) \right]^{1/2} + \frac{1+R}{1-R} \quad (5)$$

Where R is the reflectivity.

**7. Optical wavelength L:** which is the reciprocal of the absorption coefficient: [9].

$$L = 1/\alpha \quad (6)$$

**8. Optical conductivity  $\sigma_{opt}$ :** The optical conductivity is related to the refractive index and the damping factor  $k$  according to the following equation [12]:

$$\sigma_{opt} = 1/30 nku \quad (7)$$

### The Importance of the Research and its Objectives

The importance of research lies in its broad applications in scientific applications through the study and determination of the range of frequencies and transmittance of the infrared spectra of pure tin and zinc oxides and finding the transmittance, absorbance, absorption coefficient, refractive index, optical wavelength, photo-conductivity and damping factor (absorption index).

### Research Methods and Materials

#### Materials used to prepare the samples

Tin oxide (purity 99%, company TITAN BIOTECH LTD, originating in India.

Zinc oxide (purity 99%, company TITAN BIOTECH LTD, originating in India.

#### Devices and tools used

1. A sensitive balance type (SARTORIUS) with an accuracy of gr (4-10) located in the Faculty of Science - Department of Physics - Tishreen University.

2. A high-temperature convection oven (co 1200) with a temperature regulator located in the Faculty of Science - Department of Chemistry - Tishreen University.
3. Small agate mortar, Gate Morator.
4. Jasco infrared spectrometer (Plus 460-IR -TF) - Central Laboratory of the College of Science.

### Results and Discussion

FT-IR technology is used to obtain information about chemical bonds and functional groups in matter. Measuring the permeability in terms of vibrational frequency is useful in predicting the presence of some functional groups as they are absorbed at certain vibrational frequencies and so it reveals the structure of the material. Where the positions of the absorption bands and peaks depend on the crystal structure of the material and the chemical composition as well as on the morphology. FT-IR analysis was performed for the samples at room temperature within the range of (400-4000  $\text{cm}^{-1}$ ), where some absorption bands appeared for the studied samples, and we found that each sample was characterized by vibrational frequencies within the range of (400-4000  $\text{cm}^{-1}$ ).

(Tables 1-3) contain the values of vibrational frequencies  $\nu$  extracted from the IR spectra with the corresponding  $T$  values for each frequency. And refractive index  $n$ , absorbance  $A$ , damping factor  $K$ ,  $\alpha$  (absorption coefficient or linear absorption coefficient), optical wavelength  $L$ , and optical conductivity  $\sigma_{opt}$  (Figure 4) (Tables 1, 2).

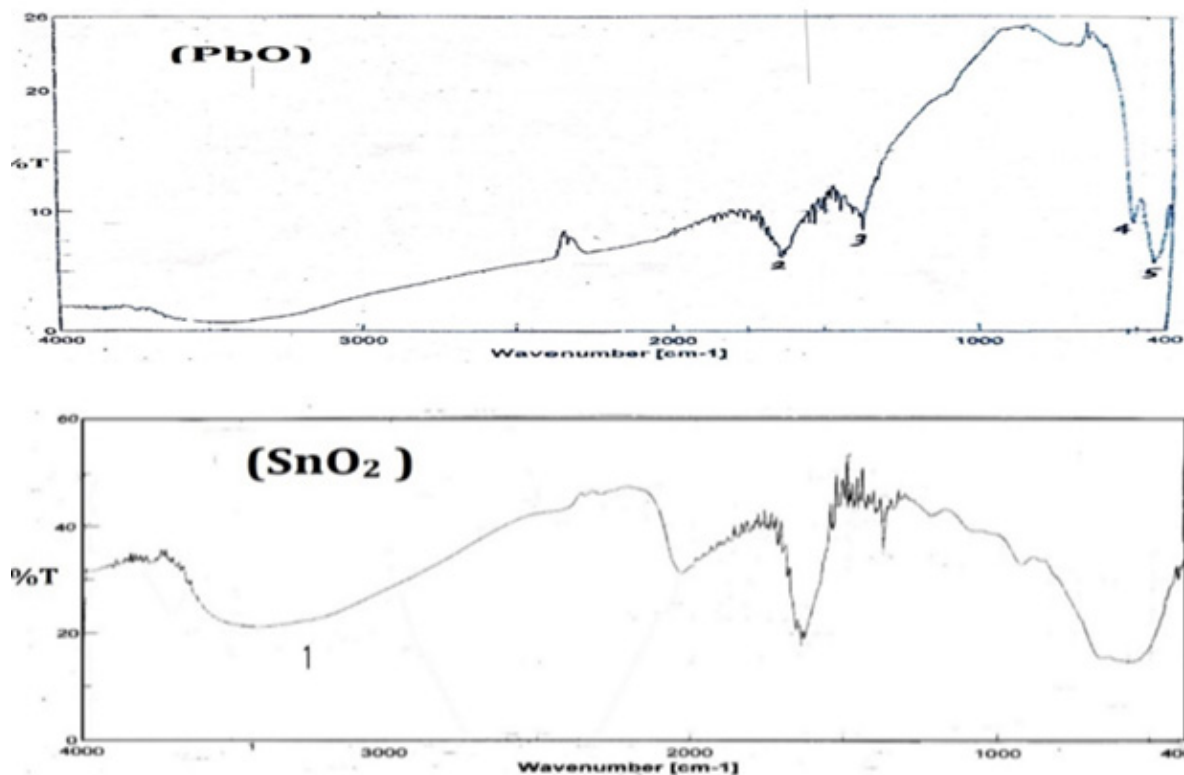


Figure 4: FTIR spectra for lead oxide and tin oxide.

**Table 1:** The table shows pure tin oxide with values of transmittance corresponding to each vibrational frequency, absorbance, reflectivity, absorption coefficient, damping factor, refractive index, optical length and optical conductivity.

$\nu$ ( $\text{cm}^{-1}$ )	T%	$n^2$	N	A	$\alpha$ ( $\text{cm}^{-1}$ )	$k \cdot 10^{-4}$	L (cm)	$\sigma_{\text{(opt)}}$	R
								( $\Omega\text{cm}^{-1}$ )	
3450	1	99	9.949	2	-1.01	46	10.62	0.021	1.214
2400	7	13.285	3.644	1.154	-0.224	26.562	8.81	0.037	0.256
1850	9.5	9.526	3.086	1.022	-0.117	23.512	10.12	0.042	0.192
1650	6.5	14.384	3.792	1.187	-0.252	27.302	13.17	0.036	0.274
1460	10	9	3	1	-0.1	23	12.54	0.043	0.183
1400	8.5	10.764	3.28	1.07	-0.92	24.623	14	0.04	0.214
750	24	3.166	1.779	0.619	0.141	14.255	15.13	0.07	0.067
425	9	10.111	3.179	1.045	-0.135	24.052	45.06	0.041	0.202
360	5.5	17.181	4.145	1.259	-0.809	28.971	64.07	0.034	0.318

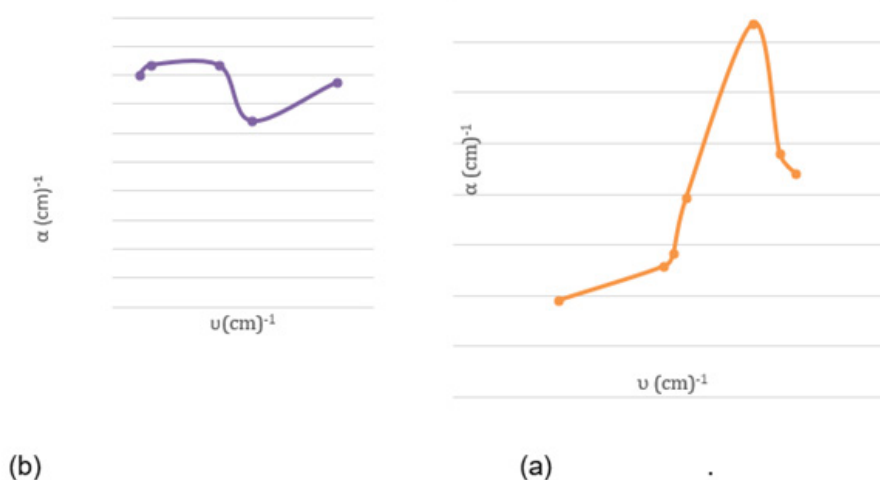
**Table 2:** The table shows pure tin oxide with values of transmittance corresponding to each vibrational frequency, absorbance, reflectivity, absorption coefficient, damping factor, refractive index, optical length and optical conductivity.

$\nu$ ( $\text{cm}^{-1}$ )	T%	$n^2$	N	A	$\alpha$ ( $\text{cm}^{-1}$ )	$k \cdot 10^{-4}$	L (cm)	$\sigma_{\text{(opt)}}$	R
								( $\Omega\text{cm}^{-1}$ )	
3432.67	21.092	4.056	2.014	0.675	15.545	3.61	0.064	0.083093	0.113
2125.3	27.692	5.621	2.371	0.557	12.825	4.8	0.077	0.080712	0.165
1641.13	18.841	3.363	1.834	0.724	16.672	8.09	0.0599	0.081178	0.0866
574.683	18.813	3.356	1.832	0.725	16.687	23.12	0.0599	0.08114	0.0863
415.585	20.188	3.786	1.946	0.694	15.982	30.62	0.062	0.082572	0.1032

[Figure (a-5)] shows the changes of the absorption coefficient in terms of the vibrational frequency of PbO, where the value of the absorption coefficient decreases with the increase of the vibrational sites until it reaches the lowest value  $\alpha = 14.255 \text{ cm}^{-1}$  for the absorption coefficient at a vibrational frequency  $\nu = 750 \text{ cm}^{-1}$ , and it begins to increase until it reaches For a value of the absorption coefficient  $\alpha = 24.623 \text{ cm}^{-1}$  at a vibrational frequency  $\nu = 1400 \text{ cm}^{-1}$ , then the value of the absorption coefficient decreases and then returns to increase, then decreases and then increases to reach the

greatest value of the absorption coefficient  $\alpha = 46 \text{ cm}^{-1}$  corresponding to a vibrational frequency of  $\nu = 3450 \text{ cm}^{-1}$ , and this is due to the bonding forces between the molecules of the compound.

[Figure (b-5)] shows the changes of the absorption coefficient in terms of the vibrational frequency of  $\text{SnO}_2$ , where the absorption coefficient increases with the increase of the vibrational sites until it reaches the highest value  $\alpha = 16.687 \text{ cm}^{-1}$ . For the lowest value  $\alpha = 12.825 \text{ cm}^{-1}$  at vibrational frequency  $\nu = 2125.30 \text{ cm}^{-1}$  (Figure 5) (Table 3).



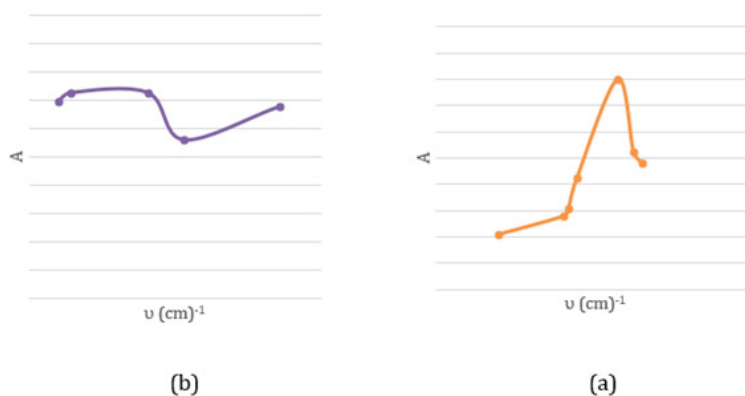
**Figure 5:** Represents changes of absorption coefficient as a function of vibrational frequency PbO (a)  $\text{SnO}_2$  (b).

**Table 3:** shows the average values of the absorption coefficient and the highest and lowest values of the absorption coefficient for the pure oxides, PbO, SnO<sub>2</sub>.

Sample	The average value of the absorption coefficient	The highest value of the absorption coefficient	The lowest value of the absorption coefficient
PbO	26.475	46	14.255
SnO <sub>2</sub>	15.542	16.687	12.825

[Figure (a-6)] shows the absorbance changes in terms of the vibrational frequency of PbO, where the absorbance decreases with the increase of the vibrational sites until it reaches the lowest value of  $A = 0.619$  for the absorbance at a vibrational frequency  $\nu = 750 \text{ cm}^{-1}$ , and it begins to increase until it reaches a value of  $A = 1.070$  at

a vibrational frequency  $\nu = 1400 \text{ cm}^{-1}$ , then the value of absorbance decreases and then returns to increase, then decreases and then increases to reach the greatest value of absorbance  $A = 2$  corresponding to a vibrational frequency of  $\nu = 3450 \text{ cm}^{-1}$ , and this is due to the weak bonding forces .

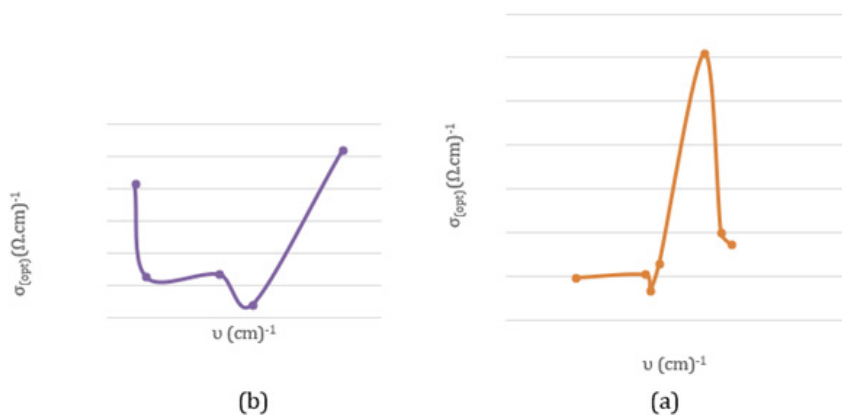


**Figure 6:** represents the absorbance changes as a consequence of the vibrational frequency SnO<sub>2</sub> (b) ZnO (a).

[Figure (a-6)] shows the changes in the absorbance in terms of the vibrational frequency of SnO<sub>2</sub>, where the absorbance increases with the increase in the vibrational sites until it reaches the highest value of  $A = 0.725$  of the absorbance at a vibrational frequency  $\nu = 574.683 \text{ cm}^{-1}$ , then decreases again until it reaches the lowest value of  $A = 0.558$  at Vibration frequency  $\nu = 2125.3 \text{ cm}^{-1}$  (Figure 6) (Table 4).

optical conductivity decreases with the increase of the vibrational sites until it reaches the lowest value  $\sigma_{opt} = 0.067323 (\Omega\text{cm})^{-1}$  of the optical conductivity at a vibrational frequency  $\nu = 750 \text{ cm}^{-1}$ , and it begins to increase Until it reaches a value of the optical conductivity  $\sigma_{opt} = 0.214406 (\Omega\text{cm})^{-1}$  at a vibrational frequency  $\nu = 1400 \text{ cm}^{-1}$ , then the value of the optical conductivity decreases and then returns to increase, then decreases and then increases to reach the greatest value of the optical conductivity  $\sigma_{opt} = 1.214687 (\Omega\text{cm})^{-1}$  corresponding to a frequency A vibrational magnitude of  $\nu = 3450 \text{ cm}^{-1}$ , due to the spatial voids between the atoms.

[Figure (a-7)] shows the changes of the optical conductivity in terms of the vibrational frequency of PbO, where the value of the



**Figure 7:** Represents the changes in optical conductivity as a consequence of the vibrational frequency SnO<sub>2</sub> (b) PbO (a).

[Figure (b-7)] shows the changes of the optical conductivity in terms of the vibrational frequency of  $\text{SnO}_2$ , where we notice from the curve the optical conductivity value decreases with the increase of the vibrational sites until it reaches the lowest value  $\sigma_{\text{opt}} = 0.0807 (\Omega\text{cm})^{-1}$  for the optical conductivity at a vibrational frequency  $\nu = 750 \text{ cm}^{-1}$ , and it starts to increase until it reaches a value of

optical conductivity  $\sigma_{\text{opt}} = 0.214406 (\Omega\text{cm})^{-1}$  at a vibrational frequency  $\nu = 2125.300 \text{ cm}^{-1}$ , and then the value of the optical conductivity increases with increasing vibrational frequency to reach the greatest value of optical conductivity  $\sigma_{\text{opt}} = 0.0831 (\Omega\text{cm})^{-1}$  corresponding to a vibrational frequency of  $\nu = 3432.67 \text{ cm}^{-1}$ , due to the spatial voids between the atoms (Figure 7) (Table 5).

**Table 4:** shows the average values of absorbance, the highest and the lowest value of absorbance for pure oxides, ZnO,  $\text{SnO}_2$ .

Sample	The average absorbance value	The lowest absorbance value	The lowest absorbance value
PbO	1.15	2	0.619
$\text{SnO}_2$	0.675	0.725	0.558

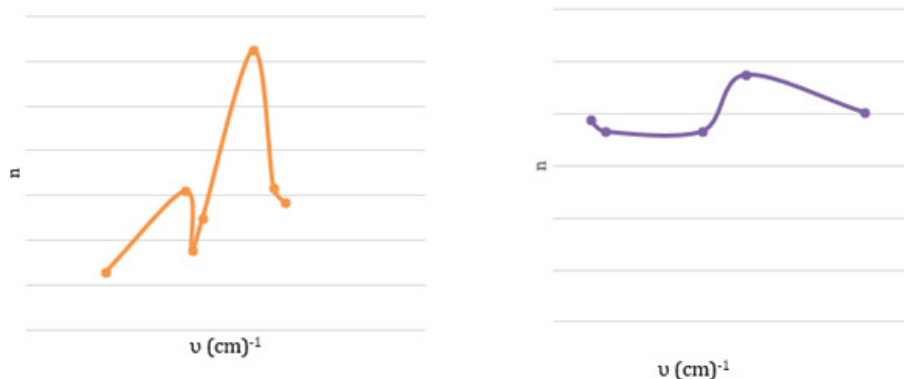
**Table 5:** shows the average values of the optical conductivity and the highest and lowest values of the photoconductivity for the pure oxides, PbO,  $\text{SnO}_2$ .

Sample	The average absorbance value	The lowest absorbance value	The lowest absorbance value
PbO	0.324	1.214687	0.067323
$\text{SnO}_2$	0.081	0.0831	0.0807

[Figure (a-8)] shows the changes of the index of refraction in terms of the vibrational frequency of PbO, where the value of the index of refraction decreases with the increase in the vibrational sites until it reaches the lowest value of  $n = 1.779$  for the index of refraction at a vibrational frequency  $\nu = 750 \text{ cm}^{-1}$ , and it begins to increase until it reaches a value for the index of refraction  $n = 3.280$  at a vibrational frequency  $\nu = 1400 \text{ cm}^{-1}$ , then the value of the refractive index decreases and then returns to increase, then decreases and then increases to reach the greatest value of the refractive index  $n = 9.949$  corresponding to a vibrational frequency of  $\nu =$

$3450 \text{ cm}^{-1}$ , and this is due to the bonding forces between compound particles.

[Figure (a-8)] shows the changes of the refractive index in terms of the vibrational frequency of  $\text{SnO}_2$ , where the refractive index decreases with the increase in the vibrational sites until it reaches the lowest value of  $n = 1.832$  the refractive index at a vibrational frequency  $\nu = 574.683 \text{ cm}^{-1}$ , then it increases again until it reaches the highest value  $n = 2.371$  at vibrational frequency  $\nu = 2125.30 \text{ cm}^{-1}$  (Figure 8) (Table 6).



**Figure 8:** represents changes of refractive index as a function of vibrational frequency  $\text{SnO}_2$  (b) PbO (a).

**Table 6:** shows the average values of the index of refraction and the highest and lowest values of the index of refraction for the pure oxides PbO,  $\text{SnO}_2$ .

Sample	The average refractive index value of the sample	The highest refractive index value	The lowest refractive index value
PbO	3.983	9.949	1.779
$\text{SnO}_2$	1.999	2.371	1.832

## Conclusions and Recommendations

1. The RI spectrum of lead oxide is characterized by nine vibrational frequencies, and tin oxide is characterized by five vibrational frequencies within the range  $(400-4000) \text{ cm}^{-1}$ .
2. The value of absorbance A for each of the studied oxides ranges within the range for lead oxide,  $[0.725-0.558]$  for tin oxide  $[0.619-2]$

3. The value of the refractive index  $n$  ranges within the range. for lead oxide, [2.371-1.832] for tin oxide [9.949-1.779].
4. The value of the optical length,  $L$ , ranges within the range for lead oxide, [0.078-0.060] for tin oxide [0.070-0.021].
5. The value of the Nuhain coefficient  $\alpha$  ranges within the range for lead oxide, [12.826-16.687] for tin oxide [14.255-46].
6. The value of the optical conductance  $\sigma$  ranges within the range for lead oxide, [0.0831-0.0807] for tin oxide [1.2146-0.06732].
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### Acknowledgement

None.

### Conflicts of Interest

No conflict of interest.

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